



Influences of temperature, power ultrasound and reaction time on the morphological properties of two new mercury(II) coordination supramolecular compounds



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ABSTRACT

Nanoparticles of two new coordination compounds, $[\text{Hg}_2(\text{L})_2(\text{Br})_4]_n$ (**1**) and $[\text{Hg}(\text{L}')(\text{SCN})_2]$ (**2**), ($\text{L} = 2\text{-amino-4-methylpyridine}$, $\text{L}' = 2,6\text{-pyridinedicarboxylic acid}$), have been synthesized by use of a sonochemical process and characterized by scanning electron microscopy (SEM), field emission scanning electron microscopy (FESEM), X-ray powder diffraction (XRPD), Fourier transform infrared spectroscopy (FTIR) spectroscopy and elemental analyses. The single crystal X-ray data of compounds **1** and **2** imply that the Hg(II) ions are four and five coordinated, respectively. Topological analysis shows that 1D and 0D coordination networks of **1** and **2** can be classified as underlying nets of topological types 2C1 and 1,3M4-1, respectively. The thermal stability of compounds has been studied by thermal gravimetric (TG) and differential thermal analyses (DTA). The role of temperature, reaction time and ultrasound irradiation power on the size and morphology of “nano-structures” **1** and **2**, has been investigated. Results claim that an increase of temperature, sonication power and decrease in reaction time leads to a decrease of particle size.

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1. Introduction

Over the last two decades, the rational design and syntheses of novel coordination supramolecular compounds are making considerable progress in the field of supramolecular chemistry and crystal engineering [1–5]. Chemists and materials scientists studied metal coordination supramolecular compounds widely, as they quite simply exhibit numerous potentially useful applications in molecular adsorption, catalysis, magnetism, luminescence, nonlinear optics, and molecular sensing which are not found in mononuclear compounds [6–11]. Mercury(II) is really a well-known metal with a variety of applications in numerous areas. However, mercury has been used for the fabrication of coordination polymers and supramolecule [12–18]. The d^{10} configuration in the Hg(II) ion is assigned to a flexible coordination environment to ensure that different geometries can be generated to tailor-made materials. The coordination polymers in this metal ion can be synthesized [13]. For almost all chemists the power of ultrasound stems from the belief that it includes a kind of energy for the modification of

chemical reactivity, which is different from that normally, used e.g. heat, light and pressure. Power ultrasound produces its effect via cavitation bubbles. These bubbles are generated over the cycle in the wave in the event the liquid structure is torn apart to create tiny voids, which collapse in the compression cycle. It has been calculated that pressures within bubbles can rise to many hundreds of atmospheres and temperature can reach several thousand degrees are degrees on the collapse of those bubbles [19]. These extreme conditions can drive chemical reactions, which have been designed to fabricate several different nano compounds [20]. In the past few years, many different types of nanomaterials have been made by this method [21–31].

Synthesis and characterization of coordination polymers with various particle sizes and morphologies are crucial to the basics of technological applications. In nano sizes' particles, the ratio of surface area to volume is vastly increased when compared to compounds with larger particle. The size effect actually starts to play a substantial role on chemical and physical properties such as catalytic activity change when compared to bulk systems [32]. The results of ultrasound radiation on chemical reactions were reported in the recent works [33–42]. In this paper, we certainly have designed a simple sonochemical process to make

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Table 1
Crystal data and structures refinement for [Hg₂(L)₂(Br)₄]_n (**1**) and [Hg(L')(SCN)₂] (**2**).

Empirical formula	C ₁₂ H ₁₆ Br ₄ Hg ₂ N ₄	C ₉ H ₅ HgN ₃ O ₄ S ₂
Formula weight	937.11	483.87
Temperature	293 (2) K	293 (2) K
Wavelength	0.71075 Å	0.71075 Å
Crystal system	Monoclinic	Triclinic
Space group	P2 ₁ /n	Pī
Unit cell dimensions	a = 7.6884 (6) Å, α = 90° b = 21.4711(16) Å, β = 95.587 (7)° c = 12.7021 (9) Å, γ = 90°	a = 8.0798 (7) Å, α = 100.489 (11)° b = 8.4352 (7) Å, β = 92.030 (12)° c = 10.5239 (9) Å, γ = 115.287 (10)°
Volume	2086.9 (3) Å ³	632.50(11) Å ³
Z	4	2
Density(calculated)	2.98247 g/cm ³	2.54051 g/cm ³
Absorption coefficient	22.354 Mg/m ³	22.354 Mg/m ³
F(0 0 0)	1664.0	448.0
Crystal size	0.37 × 0.13 × 0.1 mm ³	0.40 × 0.25 × 0.20 mm ³
Theta range for data collection	3.380–27.488°	3.716–27.575°
Reflections collected	19,816	11,708
μ	22.35 mm ⁻¹	12.51 mm ⁻¹
Index ranges	−9 ≤ h ≤ 9 −23 ≤ k ≤ 27 −15 ≤ l ≤ 16	−10 ≤ h ≤ 10 −10 ≤ k ≤ 10 −13 ≤ l ≤ 13
h, k, l (max)	9, 27, 16	10, 10, 13
(sin θ/λ) _{max}	0.649 Å ⁻¹	0.651 Å ⁻¹
Theta (max)	27.488°	27.575°
Radiation type	Mo Kα	Mo Kα
Refinement method	Full-matrix least-squares on F ²	Full-matrix least-squares on F ²
Data/restraints/parameters	4746/0/204	2902/2/176
Goodness-of-fit-on F ²	0.974	0.984
Refinement	R[F ² > 2σ(F ²)] = 0.061 wR(F ²) = 0.158 S = 0.97	R[F ² > 2σ(F ²)] = 0.037 wR(F ²) = 0.087 S = 0.95
R1 [I > 2σ(I)]	20,154, 4746, 2375	11,717, 2902, 2507
Rint	0.136	0.065
Extinction coefficient	0.00081(9)	n/a
Largest diff. peak and hole	1.710 and −1.341 e Å ⁻³	3.03, −2.35 e Å ⁻³
CCDC No.	1,442,314	1,442,310

nano-structures of [Hg₂(L)₂(Br)₄]_n and [Hg(L')(SCN)₂]. The ultrasonic power, sonicating time, the temperature of reaction were the parameters that have been changed for reaching the optimized condition. Among the potential characterization methods such as scanning electron microscopy (SEM), field emission scanning electron microscope (FESEM) and powder X-ray diffraction (PXRD) were used.

2. Experimental

2.1. Materials and physical techniques

Starting reagents for the synthesis were purchased and used without any purification from industrial suppliers (Sigma–Aldrich, Merck and others). Elemental analyses (carbon, hydrogen, and nitrogen) were performed employing a Heraeus Analytical Jena, Multi EA 3100 CHNO rapid analyzer. Fourier transform infrared spectra were recorded on a FT-IR JASCO 680-PLUS spectrometer as KBr pellets in the 4000–400 cm⁻¹ spectral range. Thermogravimetric analysis (TGA) and differential thermal analyses (DTA) of

the title compound were performed on a computer-controlled STA-PT 1500 apparatus. single-phased powder sample of 1–1, 2–1 and crystal structure of compounds **1**, **2** were loaded into alumina pans and heated with a ramp rate of 10 °C/min from room temperature to 600 °C under argon atmosphere. X-ray diffraction experiments were carried out at the MoKα wavelength at ambient temperature. A microfocused Rigaku mm003 source with

Table 2
Selected bond lengths/ Å° for compound [Hg₂(L)₂(Br)₄]_n.

Hg(1)–Br(1)	2.9542 (15)	N(4)–H(4)A	0.8623
Hg(1)–N(1)	2.233 (10)	N(4)–H(4)B	0.8495
Hg(1)–Br(3)	2.4954 (15)	O(1)W–H(1)WB	0.8500
Hg(1)–Br(2)	2.5642 (16)	C(9)–H(9)A	0.9600
Br(1)–Hg(2)	2.5622 (15)	C(13)–H(13)A	0.9600
Hg(2)–Br(2) ⁱ	2.9737 (15)	Br(2)–Hg(2) ⁱⁱ	2.9737 (15)
C(1)–N(1)	1.370 (17)	N(2)–H(2)A	0.8706
C(1)–C(2)	1.35 (2)	N(2)–H(2)B	0.8720

Symmetry transformations used to generate equivalent atoms:

ⁱ 1 + x, y, z;

ⁱⁱ −1 + x, y, z.

Table 3
Selected bond lengths/Å° for compound [Hg(L')(SCN)₂].

Hg(1)–N(1)	2.504 (5)	C(5)–C(7)	1.507 (8)
Hg(1)–S(1)	2.379 (2)	N(2)–C(9)	1.149 (9)
Hg(1)–O(1)	2.641 (4)	S(2)–C(9)	1.674 (7)
Hg(1)–S(2)	2.394(2)	C(3)–H(3)	0.9300
C(1)–N(1)	1.337 (8)	C(6)–O(4)	1.307 (8)
O(4)–H(4)A	0.8200	O(2)–C(7)	1.311 (8)
O(4)–H(4)A	0.8200	O(3)–C(6)	1.217 (9)
N(3)–C(8)	1.148 (10)	C(2)–H(2)	0.9300

Table 4
Selected bond lengths/ Å° for compound [Hg₂(L)₂(Br)₄]_n.

N1–Hg1–Br1	91.3 (3)	N3–Hg2–Br1	107.3 (3)
N1–Hg1–Br3	126.4 (3)	N3–Hg2–Br2 ⁱ	88.3 (3)
N1–Hg1–Br2	106.7 (3)	N3–Hg2–Br4	125.5 (3)
Br3–Hg1–Br1	96.01 (5)	Br4–Hg2–Br1	126.27 (6)
Br3–Hg1–Br2	125.87 (6)	Br4–Hg2–Br2 ⁱ	97.63 (5)
Br2–Hg1–Br1	92.24 (5)	Hg1–Br2–Hg2 ⁱⁱ	90.54 (4)
Hg2–Br1–Hg1	92.16 (4)	C6–N1–C1	118.0 (12)
Br1–Hg2–Br2 ⁱ	93.12 (5)	C6–N1–Hg1	123.3 (9)

Symmetry transformations used to generate equivalent atoms:

ⁱ 1 + x, y, z;

ⁱⁱ −1 + x, y, z.

integrated confocal caxFlux double bounce optic was employed, and Bragg peak measurement was performed by an HPAD Pilatus 200 K detector. The structure was solved using direct methods, refined on F^2 by means of full-matrix least-squares methods [35] and expanded using Fourier techniques. All non-hydrogen atoms were refined anisotropically. Hydrogen atoms were included in structure factor calculations from their location in difference maps. C-bound H atoms were treated as riding in geometrically idealized positions, with $U_{iso}(H) = kU_{eq}(C)$, where $k = 1.2$ for the Csp^2 – bound H atoms. For the water molecules, the oxygen–hydrogen bond lengths were restrained to 0.82(2) Å (DFIX instruction) and the hydrogen–hydrogen intramolecular distances were restrained to 1.30(4) Å (DANG instruction). The U_{iso} value for hydrogen atoms were set to 1.5 times the value of the water oxygen atom. Computing details [34–36]: data collection, cell refinement and data reduction: CrystalClear-SM expert 2.1b43 (Rigaku, 2014); program(s) used to solve structure: SHELXT [35]; program(s) used to refine structure: SHELXL-2014/7 [35]; molecular graphics: PLATON [36]; software used to prepare material for publication. A multi wave ultrasonic generator (ultrasonic homogenizer-UP 400-A, IRAN), equipped with a converter/transducer and titanium oscillator (horn), 12.5 mm in diameter, operating at 20 kHz with a maximum power output of 400 W, were used for the ultrasonic irradiation. Melting points were measured on an electrothermal 9100 apparatus and are uncorrected. X-ray powder diffraction (XRD) measurements were performed using an X'pert diffractometer manufactured by Philips with monochromatized $CuK\alpha$ radiation and simulated XRD powder patterns based on single crystal data were prepared using the Mercury software [37]. The samples were characterized with a scanning electron microscope (SEM) (Philips XL 30 and S-4160) and field emission scanning electron microscope (FE-SEM) (Perkin Elmer Elan 9000) with gold coating.

2.2. Synthesis of $[Hg_2(L)_2(Br)_4]_n$ (**1**) as single crystal

$HgCl_2$ (1 mmol, 0.271 g), 2-Amino-4-methylpyridine (1 mmol, 0.108 g) and potassium bromide (2 mmol, 0.238 g) were loaded into one arm of a branch tube and both of the arms were filled slowly by water. The chemical bearing arm was immersed in an oil bath kept at 60 °C. Crystals were formed on the inside surface of the arm kept at ambient temperature. After 10 days, colorless crystals were deposited in the cooler arm were filtered off, washed with water and air dried. (0.286 g, yield 61.12% based on final product), m.p. = 150 °C. Anal. Calc. for $C_{12}H_{16}Br_4Hg_2N_4$: C: 15.36, H: 1.70, N: 11.95%; Found C: 15.24 H: 1.33 N: 11.78 %. IR (selected bands for compound **1**; in cm^{-1}): 3454.30 (s), 3359.69 (s), 3035.49 (w), 1631.77 (vs), 1454.53 (s) cm^{-1} .

2.3. Synthesis of $[Hg_2(L)_2(Br)_4]_n$ (**1**) nano-structures under ultrasonic irradiation

To prepare the nano-structures of $[Hg_2(L)_2(Br)_4]_n$ (**1**) by sonochemical process, a high-density ultrasonic probe immersed directly into the solution of $HgCl_2$ (20 ml, 0.05 M) in water, then into this solution, a proper volume of 2-amino-4-methylpyridine ligand and KBr in water solvent (20 ml, 0.05 M) was added in a drop wise manner. The solution was irradiated by sonochemical with the power of 60 W and temperature 50 °C for 1 h. For the study of the effect of time of reaction on size and morphology of nano-structured compound **1**, the above processes were done with 30 min and for the study of the effect of sonication power with

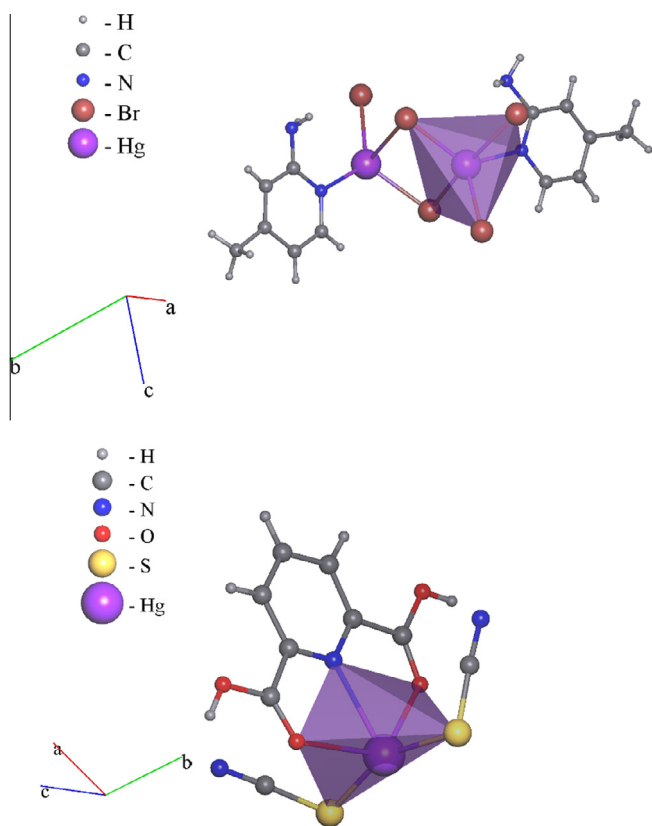


Fig. 1. The coordination environment of Hg^{2+} cation environment in compound $[Hg_2(L)_2(Br)_4]_n$ (up) and in compound $[Hg(L')(SCN)_2]$ (down).

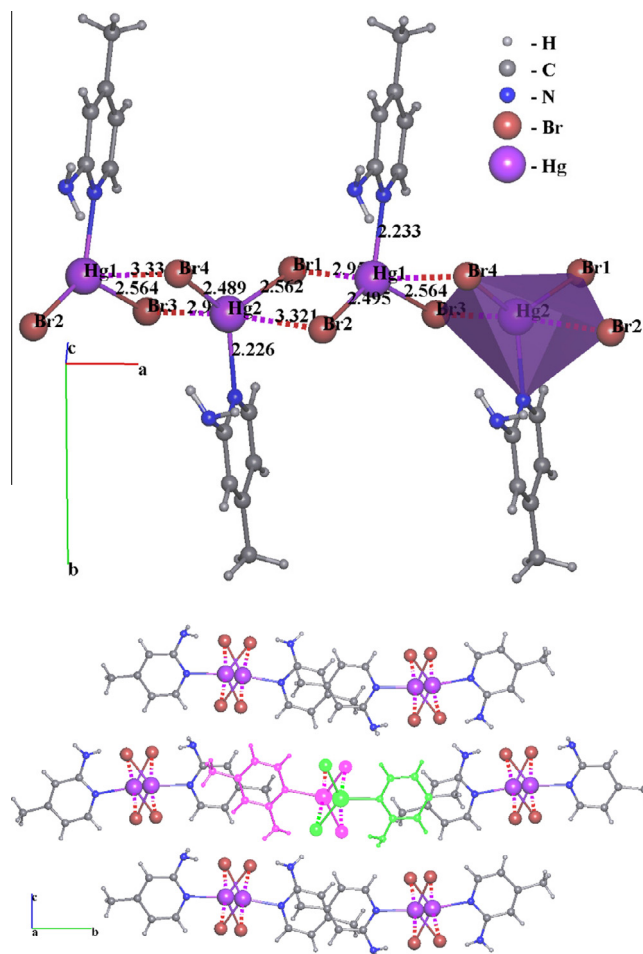


Fig. 2. Six chains surrounding the central one. Dashed lines represent van-der-Waals bonded (down). Distance of bounds in **1** (up).

80 W (time: 1 h, temperature: 50 °C, concentration: 0.05 M) also for the study of the effect of temperature with 80 °C (time: 1 h, sonication power: 60 W, concentration: 0.05 M) has been done too. The obtained precipitates were filtered, subsequently washed with water and then dried. (0.312 g, yield 64.19% based on final product), m.p. = 166 °C. Anal. Calc. for $C_{12}H_{16}Br_4Hg_2N_4$: C: 15.36, H: 1.70, N: 11.95%; Found C: 15.28 H: 1.62 N: 11.82 %. IR (selected bands for compound **1**; in cm^{-1}): 3455.48 (s), 3360.25 (s), 3035.55 (w), 1632.14 (vs), 1454.81 (s) cm^{-1} .

2.4. Synthesis of $[Hg(L')(SCN)_2]$ (**2**) as single crystal

$HgCl_2$ (1 mmol, 0.271 g), 2, 6-pyridinedicarboxylic acid (1 mmol, 0.167 g) and KSCN (2 mmol, 0.18 g) were loaded into one arm of a branch tube and both of the arms were filled slowly by water. The chemical bearing arm was immersed in an oil bath kept at 60 °C. Crystals were formed on the inside surface of the arm kept at ambient temperature, after 10 days, colorless crystals were deposited in the cooler arm were filtered off, washed with water and air dried. (0.343 g, yield 71.01% based on final product), m.p. = 180 °C. Anal. Calc. for $C_9H_5HgN_3O_4S_2$: C: 22.32, H: 1.03, N: 17.36%; Found C: 22.21 H: 0.98 N: 17.30 %. IR (selected bands for compound **2**; in cm^{-1}): 3442.27 (br), 3092.48 (w), 2109.16 (vs), 2141.45 (w), 1716.94 (vs).

2.5. Synthesis of $[Hg(L')(SCN)_2]$ (**2**) nano-structures under ultrasonic irradiation

To prepare the nano-structures of $[Hg(L')(SCN)_2]$ (**2**) by another sonochemical process, a high-density ultrasonic probe immersed

directly into the solution of $HgCl_2$ (20 ml, 0.05 M) in water, then into this solution, a proper volume of 2, 6-pyridinedicarboxylic acid ligand and KSCN in water solvent (20 ml, 0.05 M) was added in a drop wise manner. The solution was irradiated by sonochemical with the power of 60 W and temperature 50 °C for 1 h. For the study of the effect of time of reaction on size and morphology of nano-structured compound **2**, the above processes were done with 30 min and for the study of the effect of sonication power with 80 W (time: 1 h, temperature: 50 °C, concentration: 0.05 M) also for the study of the effect of temperature with 80 °C (time: 1 h, sonication power: 60 W, concentration: 0.05 M) has been done too. The obtained precipitates were filtered, subsequently washed with water and then dried. (0.356 g, yield 73.70% based on final product), m.p. ~ 199 °C. Anal. Calc. for $C_9H_5HgN_3O_4S_2$: C: 22.32, H: 1.03, N: 17.36%; Found C: 22.27 H: 1.02 N: 17.35 %. IR (selected bands for compound **2**; in cm^{-1}): 3441.55 (br), 3096.38 (w), 2108.75 (vs), 2140.93 (w), 1716.90 (vs).

3. Results and discussion

Reaction between the organic nitrogen-donor 2-amino-4-methylpyridine ligand (L), potassium bromide, and mercury(II) chloride yielded crystalline material formulated as new 1D coordination polymer $[Hg_2(L)_2(Br)_4]_n$ (**1**). Utilizing 2,6-pyridinedicarboxylic acid for N,O-donor ligand in mixture of KSCN and $HgCl_2$ leads to formation of new molecular coordination compound $[Hg(L')(SCN)_2]$ (**2**). Nanocrystals of compound **1** and **2** were obtained in aqueous solution by ultrasonic irradiation. Single crystals of compound **1** and **2**, suitable for X-ray crystallography, were prepared by thermal gradient method applied to an aqueous solution of

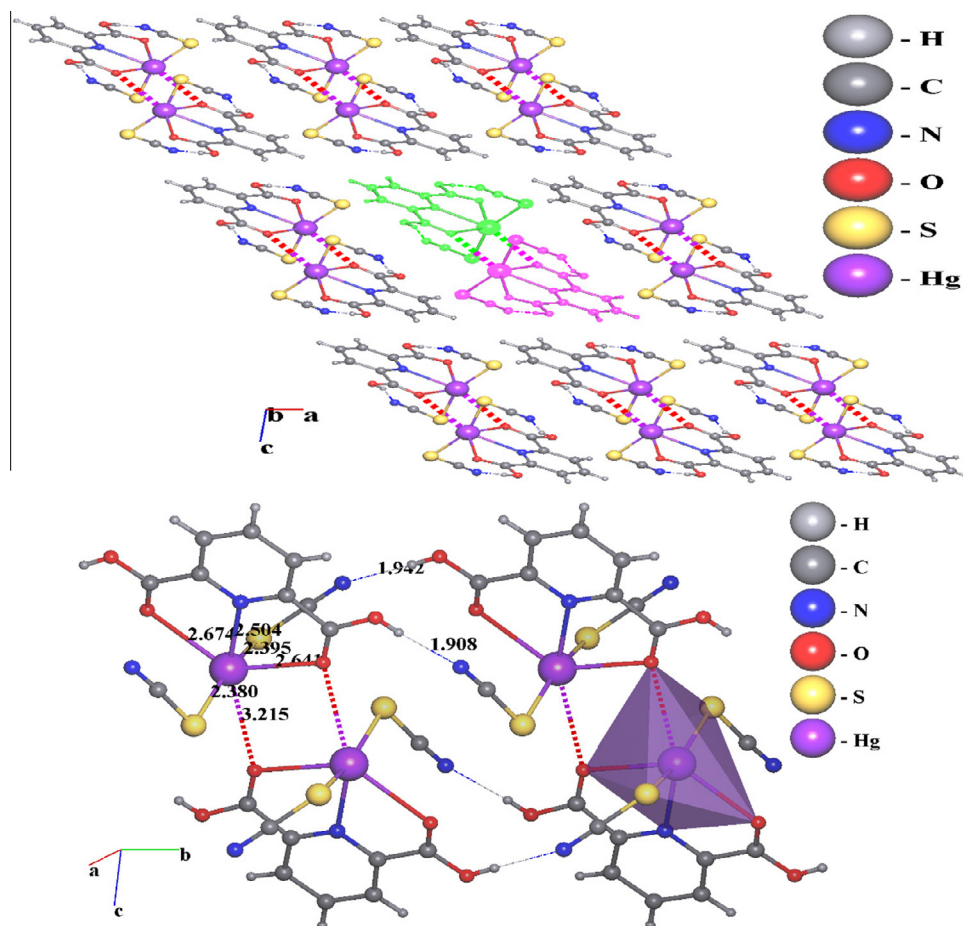


Fig. 3. Eight chains surrounding the central one. Dashed lines represent H-bonds (up). Distance of bounds in **2** (down).

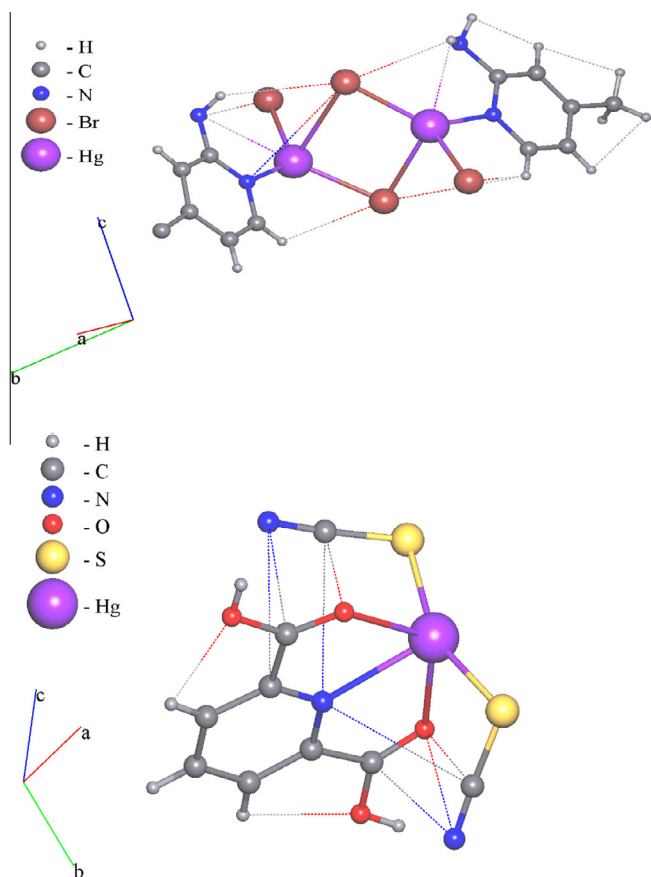


Fig. 4. Weak interaction in asymmetric units (up) compound 1, (down) compound 2.

the reagents the “branched tube method”. $[\text{Hg}_2(\text{L})_2(\text{Br})_4]_n$ and $[\text{Hg}(\text{L}')(\text{SCN})_2]$ synthesized by using the two different routes (branched tube and ultrasonic method). Single crystal X-ray diffraction analysis (Tables 1–4) of compounds **1** and **2** was carried out and the coordination environment of the title complexes are shown in (Fig. 1) Single X-ray crystal analysis reveals that $[\text{Hg}_2(\text{L})_2(\text{Br})_4]_n$ (**1**) and $[\text{Hg}(\text{L}')(\text{SCN})_2]$ (**2**) complexes crystallize in monoclinic and triclinic space group, $\text{P}2_1/\text{n}$ and $\text{P}\bar{1}$, respectively.

The Hg(II) atoms of compound **1** are coordinated by three Br atoms and one N atom composing trigonal pyramidal coordination Br_3N (Fig. 1). The asymmetric unit of compound **1** contains two Hg^{2+} cations, which coordinate two 2-amino-4-methylpyridine ligands and four Br^- anions (Fig. 4). Each 2-amino-4-methylpyridine ligand in compound **1** is coordinated to one Hg atom by N atom of pyridine ring, and Hg–N distance is about 2.23 Å. Additionally, three Br atoms are coordinated to each Hg(II) atom with contacts distances Hg–Br in the range of 2.489–2.974 Å (Tables 2 and 4 and Fig. 2). Two of four Br^- anions are bridging between two Hg^{2+} cations, and two Br^- are terminal occupying one vertex of base in trigonal pyramidal coordination polyhedron of each Hg^{2+} . The coordination interactions can be separated in two groups:

- (i). Strong, more valence, in short range of 2.226–2.564 Å;
- (ii). Weak, more electrostatic, in long range of 2.954–2.974 Å.

Both terminal Br^- ions additionally interact with neighbor Hg^{2+} ion, but with much longer distances $\text{Hg}(1)\text{--Br}(4)=3.336$ Å and $\text{Hg}(2)\text{--Br}(2)=3.321$ Å.

Strong bonds form mononuclear complexes $[\text{Hg}(\text{L})(\text{Br})_2]$, which expanded by relatively weak interactions in polymeric chain $[\text{Hg}_2(\text{L})_2(\text{Br})_4]_n$ along *a* axis (Fig. 2). Simplification of

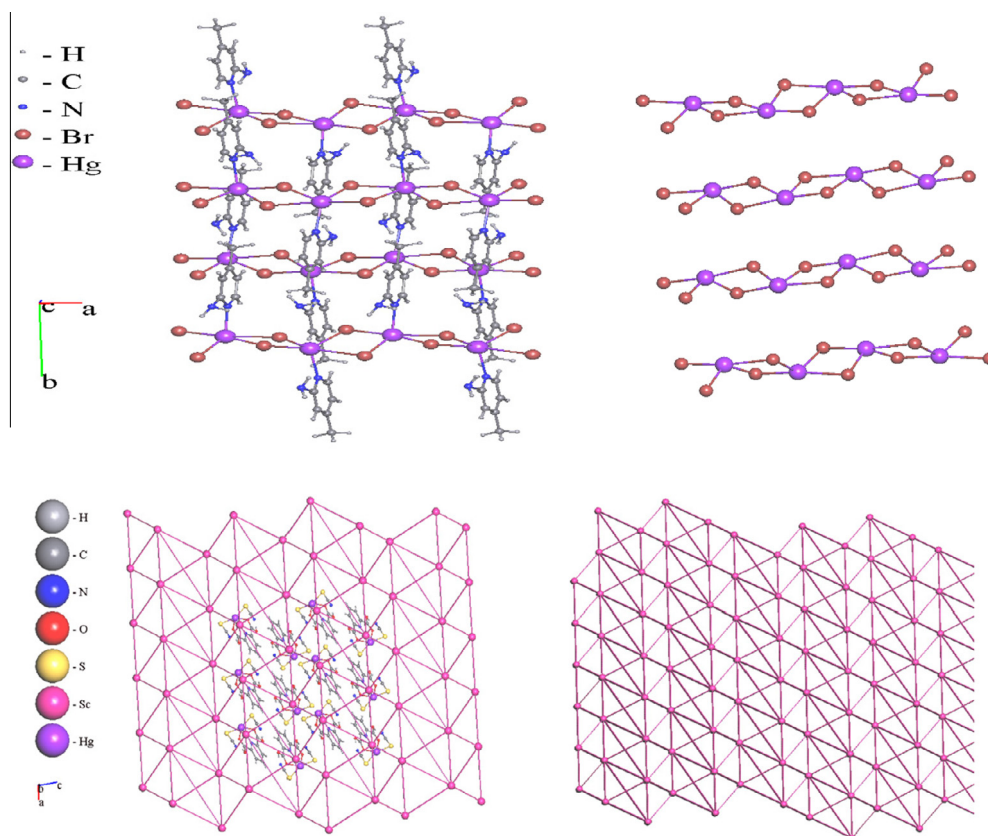


Fig. 5. Topological representation of coordination networks in compounds **1** (up) and **2** (down).

mononuclear complexes $[\text{Hg}(\text{L})(\text{Br})_2]$ into nodes of the chain underlying net and its classification by ToposPro package reveals 2C1 topological type, which is abundant for 1D coordination polymers (more than 45,000 examples in TTO collection of ToposPro (Fig. 5) [41]). Every $[\text{Hg}_2(\text{L})_2(\text{Br})_4]_n$ chain is surrounded by six other same chains forming van-der-waals bonded hexagonal packing of cylinders (Fig. 2).

In compound **2**, also Hg(II) atoms are coordinated by five atoms and have square-pyramidal coordination sphere as S_2NO_2 (Fig. 1). Asymmetric unit cell of compound **2** consists of one Hg^{2+} cation, one 2,6-pyridinedicarboxylic acid ligand, and two SCN^- anions (Fig. 4). Each Hg atom is coordinated by nitrogen (1), oxygen (2) and sulfur (2) atoms of 2,6-pyridinedicarboxylic acid ligand and SCN^- anions with Hg–N, Hg–S and Hg–O distances are about 2.380–2.674 Å (Table 3 and Fig. 3). The coordination interactions can be separated in two groups:

- (i). Relatively strong valence in range of 2.380–2.674 Å;
- (ii). More weak electrostatic Hg1–O1 = 3.215 Å.

In compound **2**, strong bonds construct monomer complex $[\text{Hg}(\text{L})(\text{SCN})_2]$, and weak interactions assemble these complexes in supramolecular dimer $[\text{Hg}(\text{L})(\text{SCN})_2]_2$ (Fig. 3) The simplification followed by the classification of the monomer complex gives 1,3M4-1 type (63,408 entries in TTO collection of ToposPro), and for dimer topological type is 1,2,4M8-3 (10,738 entries). Every $[\text{Hg}(\text{L})(\text{SCN})_2]_2$ dimer is connected to two others by H-bonds: H(4)...S(2), H(2)...N(2), H(2A)...N(3), H(4A)...N(2). The parameters of H-bonds are listed in the (Table. 6). These interactions form chains spread along *b* axis. The chains form sloping tetragonal packing of cylinders (Fig. 3).

The IR spectra display characteristic absorption bands for N–H, O–H and the 2-amino-4-methylpyridine and 2,

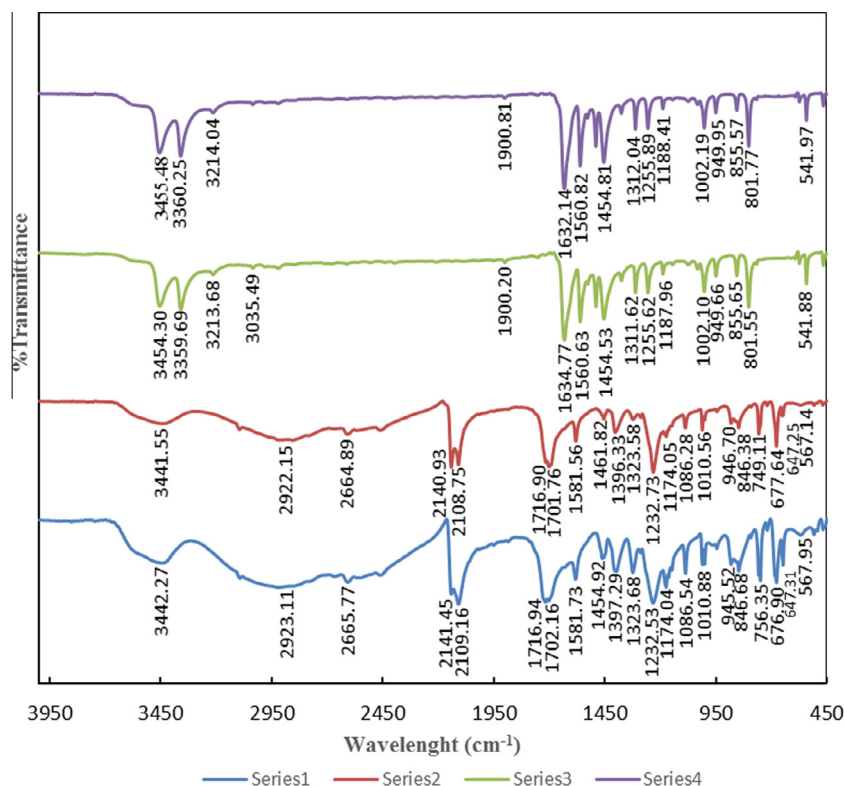


Fig. 6. The IR spectra of (Series 1) bulk materials as synthesized of compound **2**, (Series 2) nano-sized compound **2** prepared by sonochemical method, (Series 3) bulk materials as synthesized of compound **1** and (Series 4) nano-sized compound **1** prepared by sonochemical method.

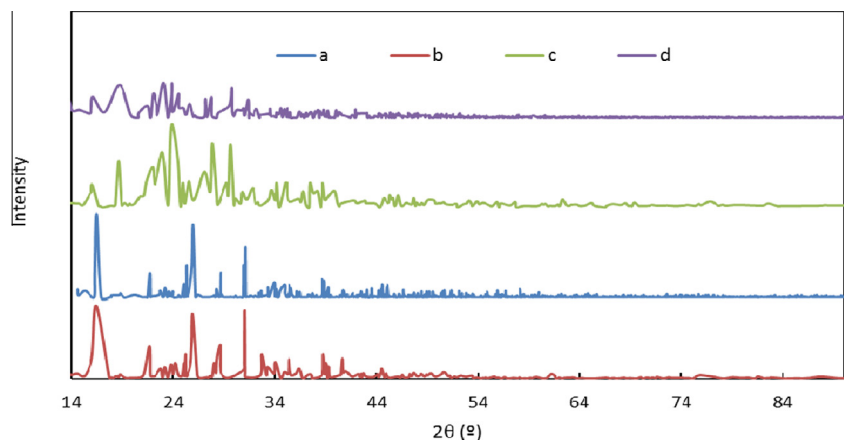


Fig. 7. The PXRD patterns of (a) nano-sized **1** obtained by ultrasonication, (b) simulated from single crystal X-ray data of **1**, (c) nano-sized **2** obtained by ultrasonication, (d) simulated from single crystal X-ray data of **2**.

6-pyridinedicarboxylic acid ligands in compounds **1** and **2**. The IR spectrum of compound **2** shows the characteristic stretching frequency of O–H group observed at about 3442 cm^{-1} the relatively broad absorption bands, The absorption bands with variable intensity in the frequency range $1700\text{--}1750\text{ cm}^{-1}$ correspond to C=O stretching frequency from the carbonyl of the 2, 6-pyridinedicarboxylic acid ligands. Also, characteristic band of the C–O stretching frequency carboxylic acid group appears at $1000\text{--}1300\text{ cm}^{-1}$. The absorption bands around 2141 and 2109 cm^{-1} related to stretching frequency of the SCN group. For compound **1** two bands around 3454 and 3359 cm^{-1} are due to the N–H stretching frequency amine group ligand. The absorption bands with variable intensity in the frequency range $1640\text{--}1560\text{ cm}^{-1}$ correspond to N–H bending frequency of the amine group of the 2-amino-4-methylpyridine ligand, absorption bands around 1454 cm^{-1} related to CH_3 bending. Also, aromatic C–H stretching frequency appears at around 3035 cm^{-1} . The elemental analysis and IR spectra of the nano-structure produced by the sonochemical method as well as the bulk material produced by the branched tube method are indistinguishable (Fig. 6).

Fig. 7b and d shows the XRD patterns of compounds **1** and **2** simulated from single crystal X-ray data. While the experimental XRD patterns of compounds **1** and **2** prepared by the sonochemical process is shown in Fig. 7a and c, respectively. For the both

compounds, acceptable matches, with slight differences in 2θ , were observed between the simulated and experimental powder X-ray diffraction patterns. This indicates that the compound **1** and **2** obtained by the sonochemical process as nano-structures are identical to that obtained by single crystal diffraction. The significant broadening of the peaks indicates that the particles are of nanometer dimensions. As Fig. 7 shows the peak is sharper and then crystal lattice will be more regular. Also, for crystal and nanostructures of compounds **1** and **2** are concluded which have same phase.

In order to look at the thermal stability of the two compounds, thermal gravimetric (TG) and differential thermal analyses (DTA) were carried out for compounds **1** and **2** between 30 and $600\text{ }^\circ\text{C}$ under argon flow. The TG curve of compound **1** indicates that the compound does not melt and is stable up to $100\text{ }^\circ\text{C}$ at that temperature it begins to decompose (Fig. 8). Removal of the 2-amino-4-methylpyridine ligand connected to Hg atom takes place in the range between 190 and $330\text{ }^\circ\text{C}$ with two exothermic effects at $199.87\text{ }^\circ\text{C}$ and between 250 and $300\text{ }^\circ\text{C}$ with a mass loss of 99.82% , ultimately giving amorphous solid that appears to be HgBr_2 . The TG curve of compound **2** indicates that this compound is stable up to $190\text{ }^\circ\text{C}$, at which temperature it begins to decompose (Fig. 8). Removal of the three carboxylate anions, one aromatic ring of 2,6-pyridinedicarboxylic acid ligand, and decomposition of the

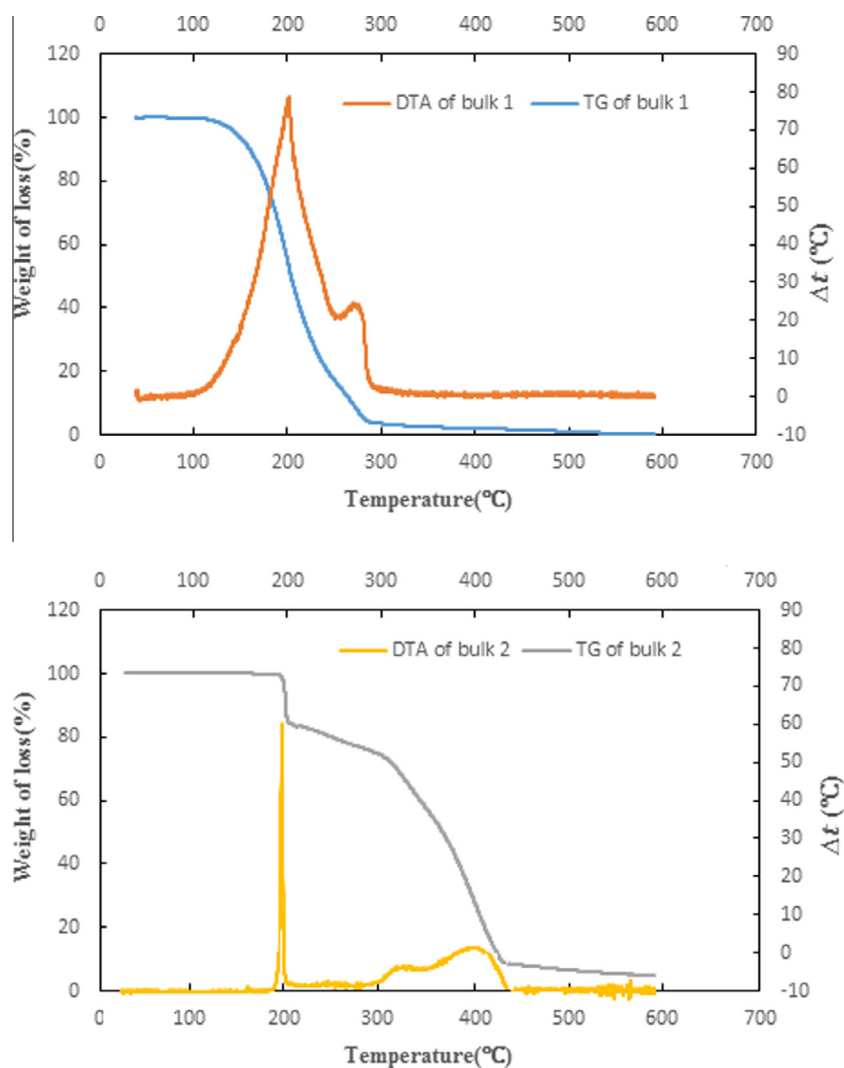


Fig. 8. Thermal behavior of compound **1** (up) and **2** (down).

Table 5

The influence of temperature, reaction time and sonication power on the size of compound **1** and **2** particles.

Compound 1 samples	T (°C) ^a	T (min) ^b	Sonication (input power) (W)	Size ^c (nm)
1-1	50	60	0	617
1-2	50	60	60	79
1-3	50	30	60	55
1-4	70	60	60	42
Compound 2 samples	T (°C) ^a	T (min) ^b	Sonication (input power) (W)	Size ^c (nm)
2-1	50	60	0	590
2-2	50	60	60	85
2-3	50	30	60	57
2-4	70	60	60	53

^a Reaction temperature.

^b Reaction time.

^c Average diameter (nm).

thiocyanate anions takes place in three steps between 190 and 450 °C and mass loss of 95.13% with three exothermic between 190 and 550 °C.

Various conditions for preparation of compounds **1** and **2** nanostructures were summarized in Table 5. In this table, sample 1-1 and 2-1 were studied without power ultrasound and the other samples were studied under variable temperature, time and power ultrasound. In order to research the role of power ultrasound irradiation on the character of product, reactions were performed under completely different power ultrasound irradiation. Results show a decrease in the particles size as increasing power ultrasound irradiation [38–40].

In sample 1-1, the reactions were studied without power ultrasound. Results show that size particles sample of 1-1

(Fig. 9a) is larger than 1-2 and 1-4 (Fig. 9b), these results are similar with samples 2-1, 2-2 and 2-4 (Fig. 10a, b). Table 6, shows the average diameter field emission scanning electron microscope (FESEM) and scanning electron micrographs (SEM) of the prepared samples. Results show high power ultrasound irradiation decreased agglomeration, and thus led to decrease particles size. Comparison between samples 1-2 and 1-3 shows a decrease in nanoparticle size. Thus, size particles of sample 1-3 are smaller than 1-2 (Fig. 9b, c). These facts are repeated in samples 2-2 and 2-3 (Fig. 10b, c). Table 6, shows the average diameter of particles shown by field emission scanning electron microscope (FESEM) and scanning electron micrographs (SEM) of the prepared samples. However, a reducing the reaction time led to the decrease of size particles of sample 1-3. Particle sizes and morphology of nanoparticle are depending on temperature [38]. Higher temperature (70 °C) results in an increased solubility, and thus a reduced supersaturation of growth species in the solution, and thus particles size of sample 1-4 is smaller than particles size of sample 1-2 (Fig. 9d, b). These facts are repeated in samples 2-4 and 2-2 (Fig. 10d, b). Table 5, shows the average diameter of particles shown by field emission scanning electron microscope (FE-SEM) and scanning electron micrographs (SEM) of the prepared samples. The most important point about **2**, there are different morphology between Figs. 10(a, b) and 10(b, c). Pay attention to, the best conditions for getting a small sized and less agglomerated nanostructure materials for these two mercury coordination supramolecular compounds are temperature, reaction time and the power of ultrasonic irradiation 70 °C, 60 min and 60 W, respectively. Also, other condition (temperature = 50 °C, reaction time = 30 min and power of ultrasonic irradiation = 60 W) is good for obtaining small nanoparticle size. These two conditions led to different morphology in **2**.

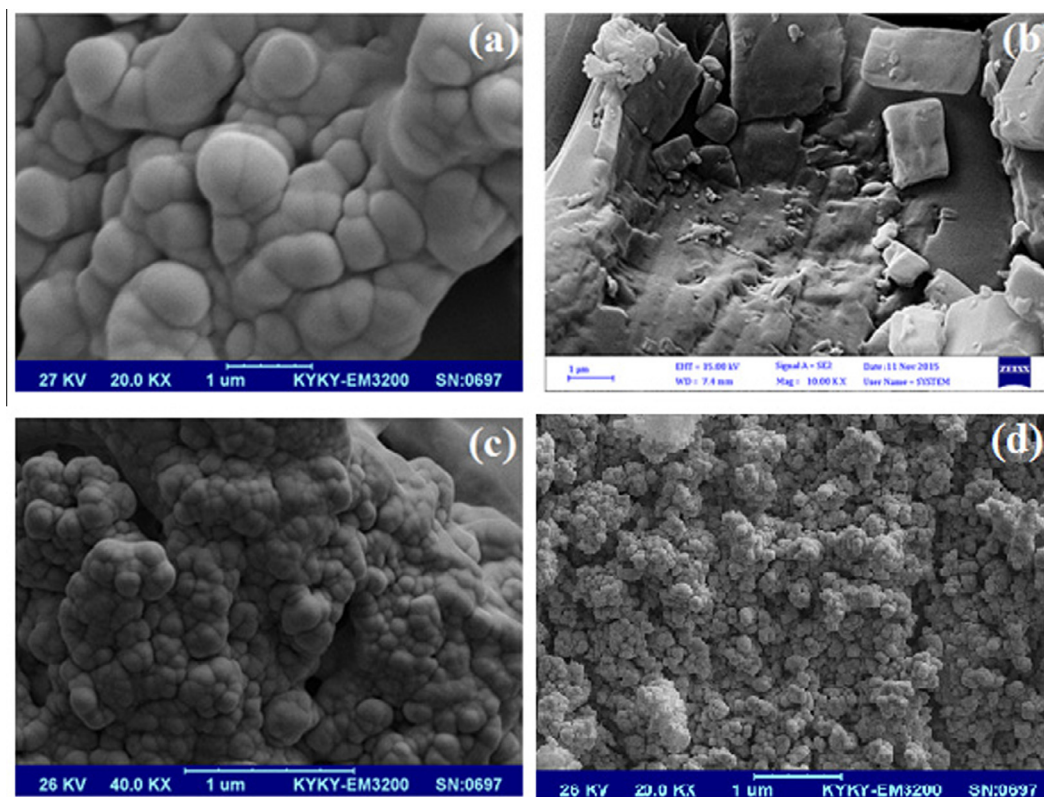


Fig. 9. FE-SEM and SEM image of nano particle **1**, (a) without sonochemical reaction, (b) by sonochemical reaction with sonochemical temperature of 50 °C, time of 60 min and power of 60 W, (c) by sonochemical reaction with sonochemical temperature of 50 °C, time of 30 min and power of 60 W, (d) by sonochemical reaction with sonochemical temperature of 70 °C, time of 60 min and power of 60 W.

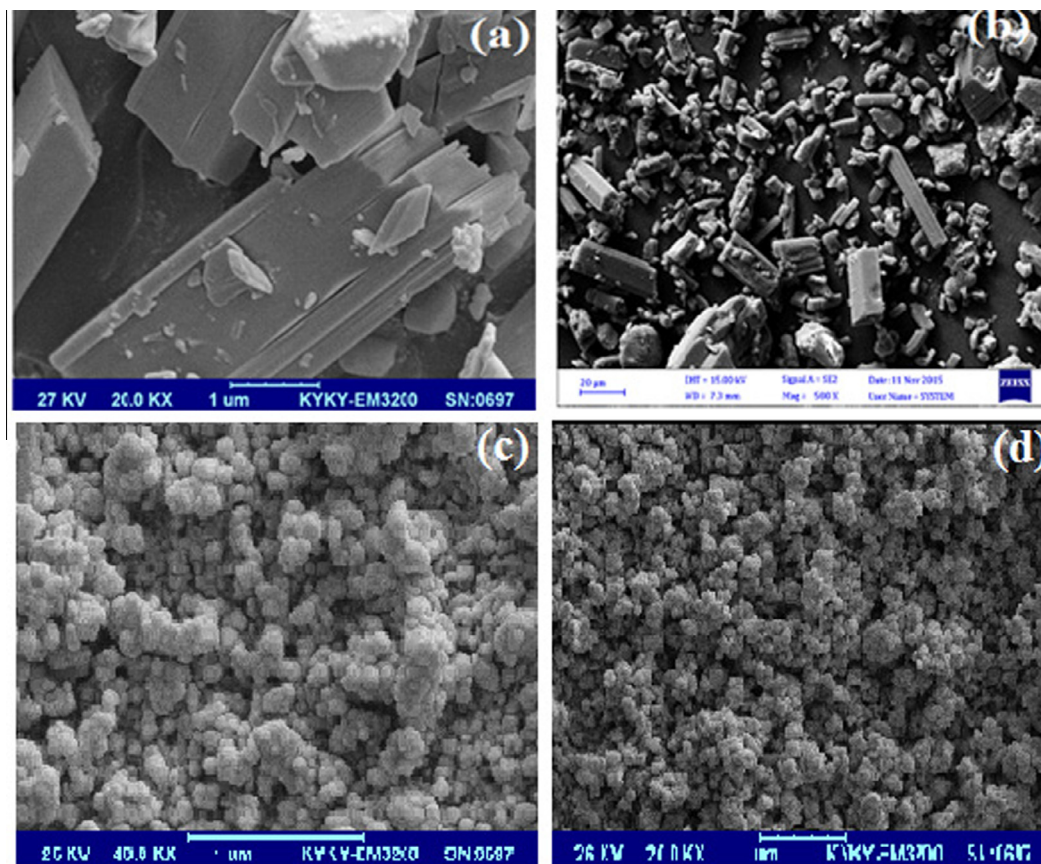


Fig. 10. FE-SEM and SEM image of nano particle 2, (a) without sonochemical reaction, (b) by sonochemical reaction with sonochemical temperature of 50 °C, time of 60 min and power of 60 W, (c) by sonochemical reaction with sonochemical temperature of 50 °C, time of 30 min and power of 60 W, (d) by sonochemical reaction with sonochemical temperature of 70 °C, time of 60 min and power of 60 W.

Table 6
Hydrogen bond details, distances (Å) and angles (°) for [Hg(L')(SCN)₂].

D–H...A	d (D–H)	d (H...A)	d (D...A)	<(DHA)
C(4)–H(4)...S(2) ^{#1}	0.93	3.01	3.900 (7)	160.7
C(2)–H(2)...N(2) ^{#2}	0.93	2.54	3.404 (9)	154.0
O(2)–H(2A)...N(3) ^{#3}	0.831 (18)	1.904 (18)	2.712 (8)	164 (7)
O(4)–H(4A)...N(2) ^{#4}	0.82	1.94	2.757 (8)	172.7

Symmetry transformations used to generate equivalent atoms:

- ^{#1} x+1,y,z.
- ^{#2} -x+1,-y+1,-z+2.
- ^{#3} x,y-1,z.
- ^{#4} x,y+1,z.

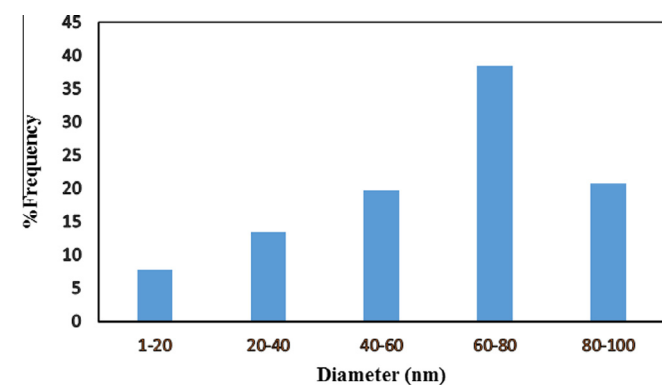


Fig. 11. Corresponding particle size distribution histogram of compound 1 nanoparticles prepared by ultrasonic generator 60 W in concentration of initial reagents [Hg²⁺] = [L⁻] = [Br⁻] = 0.05 M.

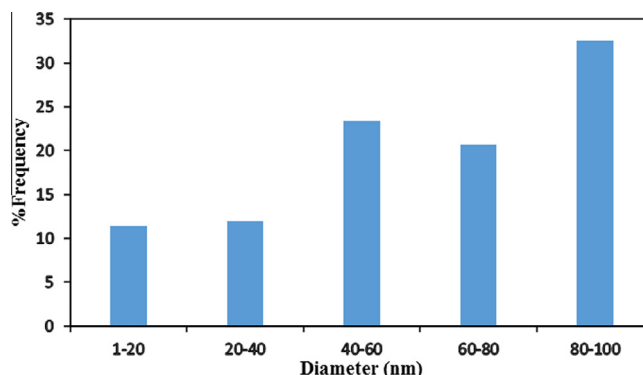


Fig. 12. Corresponding particle size distribution histogram of compound 2 nanoparticles prepared by ultrasonic generator 60 W in concentration of initial reagents [Hg²⁺] = [L⁻] = [SCN⁻] = 0.05 M.

Figs. 11 and 12 shows corresponding particle size distribution histogram of compound 1 and 2 nanoparticles prepared by ultrasonic generator 60 W in concentration of initial reagents of 0.05 M too. In addition, is showed which maximum of frequency 1 and 2 are located range of 60–80 nm and 80–100 nm, respectively.

4. Conclusion

Two new Hg(I) coordination supramolecular compound [Hg₂(L)₂(Br)₄]_n (1) and [Hg(L')(SCN)₂] (2), (L = 2-amino-4-

methylpyridine, $L' = 2,6$ -pyridinedicarboxylic acid), have been synthesized utilizing a thermal gradient approach and also by sonochemical irradiation. Compounds **1** and **2** were structurally characterized by single crystal X-ray diffraction. The crystal structures of compound **1** and **2** are made up of two 1D and 0D supramolecular compounds and show the relevant coordination number for the Hg(II) ions is found to be four and five, respectively. Influences of temperature, power ultrasound and reaction time on the morphological properties of $[Hg_2(L)_2(Br)_4]_n$ and $[Hg(L')(SCN)_2]_n$, were studied. These parameters have noticeable influences in the morphology of the nanoparticles. These systems depicted a decrease in the particles size accompanying an increase of the temperature, and of the power ultrasound as well as a reduction in time reaction. Also, the best conditions for getting small sized and less agglomerated nanostructure materials for these two mercury coordination supramolecular compounds are temperature, reaction time and the power of ultrasonic irradiation 70 °C, 60 min and 60 W, respectively.

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